

## Solutions Worksheet 1 Molarity Complete The Table Answers

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### Solutions Worksheet 1 Molarity Complete

Solutions Worksheet #1: Molarity 1) Complete the table: Solute formula Molar Mass of solute (g/mole) Mass of Solute (g) Moles of solute (mole) Molarity of solution (M) Volume of solution (L) a NaCl 1.0 1.0 b NaOH 117.0 4.0 c MgCl<sub>2</sub> 190.3 2.0 d NaCl 292.5 0.5 e KBr 238.0 2.0 f NaCl 1.5 1.0 g NaCl 1.0 0.25 h FeCl<sub>3</sub> 0.1 0.50 i HCl 438.0 2.0 j NH<sub>3</sub> 39.1 3.0

### Solutions Worksheet #1: Molarity Molar Mass of Mass of ...

Name \_\_\_\_\_ Date \_\_\_\_\_ Period \_\_\_\_\_ Solutions Worksheet #1: Molarity 1) Complete the table: Solute formula Molar Mass of solute (g/mole) Mass of Solute (g) Moles of solute (mole) Molarity of solution (M) Volume of solution (L) a NaCl 1.0 1.0 b NaOH 117.0 4.0 c MgCl<sub>2</sub> 190.3 2.0 d NaCl 292.5 0.5 e KBr 238.0 2.0 f NaCl 1.5 1.0 g NaCl 1.0 0.25 h ...

### molarconc.pdf - Name Date Period Solutions Worksheet#1 ...

NaCl in enough water to make 200.0 mL of solution. 4. Calculate the volume of 3.00 M copper(II) sulfate that must be diluted with water to prepare 500.0 mL of 2.15 M copper(II) sulfate. 5. Outline the correct laboratory procedure for diluting a 15.8 M stock solution of hydrochloric acid to a 1.0 L of a 3.0 M solution..00625 mol MgSO<sub>4</sub> (120g MgSO<sub>4</sub>/1 mol MgSO<sub>4</sub>) = .752 g MgSO<sub>4</sub>,250L (.025 mol MgSO<sub>4</sub> ...

### Molarity and Solutions WS complete.pdf - Name&Date AP ...

Name: Date: Molarity About Chemistry <http://chemistry.about.com> Complete the table for the following aqueous solutions

### Name: Date: Molarity

Molarity Worksheet W 331 Everett Community College Student Support Services Program What is the molarity of the following solutions given that: 1) 1.0 moles of potassium fluoride is dissolved to make 0.10 L of solution. 2) 1.0 grams of potassium fluoride is dissolved to make 0.10 L of solution.

### Molarity Worksheet W 331 - Everett Community College

5. 125 cm<sup>3</sup> of solution contains 3.5 moles of solute. What is the molarity of the solution? ? g KNO<sub>3</sub> = 0.175 mol KNO<sub>3</sub> × 101.1 g KNO<sub>3</sub> 1 mol KNO<sub>3</sub> = 17.7 g KNO<sub>3</sub> M = 3.5 mol 0.125 L = 28 M 6. Which solution is more concentrated? Solution "A" contains 50.0 g of CaCO<sub>3</sub> in 500.0 mL of solution. Solution "B" contains 6.0 moles of H<sub>2</sub>SO<sub>4</sub> ...

### Molarity: Molarity = 1. 2. - Central Bucks School District

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### Solutions Worksheet 1 Molarity Complete The Table Answers

Molarity Problems Worksheet M=nV n= # moles V must be in liters (change if necessary) 1. What is the molarity of a 0.30 liter solution containing 0.50 moles of NaCl? 2. Calculate the molarity of 0.289 moles of FeCl<sub>3</sub> dissolved in 120 ml of solution? 3. If a 0.075 liter solution c...

### Molarity and Dilutions Worksheet - Google Docs

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### Solutions Worksheet 1 Molarity Complete The Table Answers

Key+ 1)++23.5g+of+NaCl+is+dissolved+in+enough+water+to+make+.683L+of+solution. + a)++What+is+the+molarity+(M)+of+the+solution?+ ++ Molar+mass+of+NaCl+=58.44g/mole+ Moles+of+NaCl+ ...

### Calculations+for+Solutions+Worksheet+and+Key+

M 1 V 1 = M 2 V 2 (1.71 M)(25.0 mL) = M 2 (65.0 mL) M 2 = 0.658 M; M = mol/L = (25.0/40.0) / (0.325) = 1.92 mol/L; g = (M)(L)(FW) = (0.400)((0.225)(119) = 10.7 g (25.0g)(1 mol/101 g)(1000mL/0.650 mol) = 381 mL; Zn(NO 3) 2 AlCl 3 CuAc 2

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### Solutions Worksheet 1 Molarity Complete The Table

Solutions Homework 1. Find the molarity of a sol'n made by dissolving 44.2 g of ammonium sulfate, (NH 4) 2 SO 4, in 600.0 mL of sol'n. 0.557 M 2. What mass (in grams) of calcium nitrate, Ca(NO 3) 2, is present in 400.0 mL of a 0.150 M solution? 9.85 g. 3. What volume (in mL) of a 3.00 M solution would contain 6.50 g of lithium silicate, Li

### Solutions - Molarity, Molality, and Dilutions

Chromatography Lab Worksheet. Chromatography is a way to look at complex mixtures by separating them into their components. Criminal investigators use this technology to identify substances such as drugs, blood, ink and other fluids. Here's how it works: Different inks have different properties, such as how much they can be dissolved in solvents.