

Chapter 11

Stoichiometry Study Guide Answer Key

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Chapter 11 study guide True/False

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Indicate whether the statement is true or false. ____ 1. The actual yield is always lower than the theoretical yield.

Matching Match each item with the correct statement below. a.

stoichiometry ...

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CHAPTER Section 11.3 Limiting Reactants Date Class STUDV GUIDE In your textbook, read about why reactions stop and how to determine the limiting reactant. Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 25–27

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stoichiometry.

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Section 11.1 What is stoichiometry? In your textbook, read about stoichiometry and the balanced equation. For each statement below, write true or false. 2. 3. 4. 5. The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry. Stoichiometry is based on the law of conservation of mass,

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In Section 11.3 , for example, you learned how to express the stoichiometry of the reaction for the ammonium dichromate volcano in terms of the atoms, ions, or molecules involved and the numbers of moles, grams, and formula units of each (recognizing, for instance, that 1 mol of ammonium dichromate produces 4 mol of water). This section ...

Chapter 11: Stoichiometry

Stoichiometry Section 11.1 What is stoichiometry? In your textbook, read about stoichiometry and the balanced equation. For each statement below, write true or false. ____ 1. The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry.

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11.1 Defining Stoichiometry MAIN Idea

The amount of each reactant present at the start of a chemical reaction

determines how much product can form.

11.2 Stoichiometric Calculations MAIN

Idea The solution to every stoichiometric problem requires a balanced chemical

equation. 11.3 Limiting Reactants MAIN

Idea A chemical reaction

Chapter 11: Stoichiometry

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Stoichiometry (Rough outline of the chapter, please use the book, notes & homework to study.) 11.1 Defining

Stoichiometry Vocab • stoichiometry •

mole ratio Concepts Using Balanced

Equations • Number of Atoms • Number

of Molecules • Number of Moles • Mass o

Law of Conservation of Mass • Volume.

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CHAPTER Section 11.1 continued In your textbook, read about mole ratios.

Answer the questions about the following chemical reaction. sodium + iron(III) oxide

\rightarrow sodium oxide + iron

$6\text{Na(s)} + 2\text{Fe}_2\text{O}_3\text{(s)} \rightarrow 3\text{Na}_2\text{O(s)} + 4\text{Fe(s)}$

15. What is a mole ratio? ... Study Guide Chemistry: Matter and Change Chapter 11 ...

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Name CHAPTER Section 11.2 continued

Date Class STUDY GUIDE In your

textbook, read about mole-to-mass and

mass-to-mass conversions. Solving a

mass-to-mass problem requires the four

steps listed below.

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11 YCHAPTER Section 11.2

Stoichiometric Calculations In your

textbook, read about mole-to-mole

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conversion. Read the following passage and then solve the problems. In the equation that follows ... Chemistry: Matter and Change Chapter 11 o. Iso Study Guide 74 ...

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The Stoichiometry chapter of this Glencoe Chemistry - Matter and Change companion course helps students learn the essential chemistry lessons of products and reactants.

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